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 $K = \frac{[N_2]^{1/2} [H_2O]^2}{[NH_3]^2}$  Use of the Lewis and Randall rule gives:  $f_i = y_i P$  at 1 atm.

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POWER LAW BASICS | 6 2 POWER LAW BASICS 2.1

ARRHENIUS EQUATION The main assumption behind the

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Arrhenius expression is that  $k = A \exp(-E_a/RT)$ . This is an approximation, but it works quite well. The rate coefficient is the term that is a function of temperature but may also depend

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Chemical Equilibrium  $dG = -SdT + VdP + \sum_{j=1}^n \mu_j dn_j$   
chemical potential for species  $j$ .  $G_i(T,P) = \sum_{j=1}^n \mu_j$   
 $0, i = 1, \dots, n$   
 $\mu_j = G_j + RT \ln a_j, a_j = f_j/f_j^K$   
 $i = \sum_{j=1}^n \nu_{ij} a_{ij}$   
 $G_i = -RT \ln K_i$  Standard state: pure species  $j$  at 1 atm and system temperature.

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Description. Chemical kinetics is the study of the rates of chemical reactions; if changes in conditions impact the speed of a reaction, we can better understand what caused the reaction. In this course, you ' ll learn to understand complicated data sets and analysis techniques for measuring and understanding these rate changes.

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