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Write correct formulas for the products in these double replacement reactions.  $\text{H}_3\text{PO}_4$   
2)  $\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow$  3)  $\text{Al} + \text{HNO}_3 \rightarrow$  4)  $\text{AgNO}_3 + \text{KCl} \rightarrow$  5)  $\text{H}_2\text{SO}_4 + \text{Co} \rightarrow$  6)  $\text{BCl}_2 + \text{H}_2\text{O} \rightarrow$  7)  $\text{P}_4\text{O}_{10} + \text{CaS} \rightarrow$  8)  $\text{Ag}_2\text{CO}_3 + \text{K}_2\text{CrO}_4$

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LLabab What evidence can you observe that a reaction is taking place? During a chemical reaction, reactants are consumed as new products are formed. Often, there are several telltale signs that a chemical reaction is taking place. Procedure 1.

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As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano. Heat, light, and gas are produced as a large pile of fluffy green chromium(III) oxide forms. We can describe this reaction with a chemical equation An expression that gives the identities and quantities of the ...

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